

## Chapter 34 Problem 65 †

### Given

$$q = 2e$$

### Solution

a) Find the radius of the ground-state electron in a singly ionized Helium atom,  $He^-$ .

The radius of the ground-state can be found using the Bohr radius

$$a_0 = \frac{\hbar^2}{mke^2} = \frac{(h/2\pi)^2}{mke^2} = \frac{h^2}{4\pi^2mke^2}$$

The value of  $e^2$  in the denominator is due to the charge of the electron and the charge of the nucleus. Since the nucleus has twice the charge, we need to replace  $e^2$  with  $2e^2$ .

$$a_0 = \frac{h^2}{4\pi^2mk2e^2} = \frac{h^2}{8\pi^2mke^2}$$

Substituting in the appropriate values gives

$$a_0 = \frac{(6.63 \times 10^{-34} \text{ J} \cdot \text{s})^2}{8\pi^2(9.11 \times 10^{-31} \text{ kg})(8.99 \times 10^9 \text{ Nm}^2/\text{C}^2)(1.6 \times 10^{-19} \text{ C})^2} = 2.66 \times 10^{-11} \text{ m} = 26.6 \text{ pm}$$

b) Find the photon energy emitted when the single electron makes a transition from the  $n = 2$  to the  $n = 1$  state.

The energy level of each state is given by the equation

$$E = -\frac{ke^2}{2a_0} \left( \frac{1}{n^2} \right)$$

Since the nuclear charge is doubled, the value of  $e^2$  in the numerator should go to  $2e^2$  as discussed in part a. This gives

$$E = -\frac{k2e^2}{2a_0} \left( \frac{1}{n^2} \right) = -\frac{ke^2}{a_0} \left( \frac{1}{n^2} \right)$$

The energy for state  $n = 1$  is

$$E_1 = -\frac{(8.99 \times 10^9 \text{ Nm}^2/\text{C}^2)(1.6 \times 10^{-19} \text{ C})^2}{2.66 \times 10^{-11} \text{ m}} \left( \frac{1}{1^2} \right) = -8.65 \times 10^{-18} \text{ J}$$

The energy for state  $n = 2$  is

$$E_2 = -\frac{(8.99 \times 10^9 \text{ Nm}^2/\text{C}^2)(1.6 \times 10^{-19} \text{ C})^2}{2.66 \times 10^{-11} \text{ m}} \left( \frac{1}{2^2} \right) = -2.16 \times 10^{-18} \text{ J}$$

The change in energy is then

$$\Delta E = E_1 - E_2 = -8.65 \times 10^{-18} \text{ J} - (-2.16 \times 10^{-18} \text{ J}) = -6.49 \times 10^{-18} \text{ J}$$

An energy loss by the atom must be transmitted away as a photon. Therefore, the energy of the photon is  $6.49 \times 10^{-18} \text{ J}$ . Converting this to electron volts gives us

$$\Delta E = (6.49 \times 10^{-18} \text{ J}) \left( \frac{1 \text{ eV}}{1.6 \times 10^{-19} \text{ J}} \right) = 40.6 \text{ eV}$$

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†Problem from Essential University Physics, Wolfson